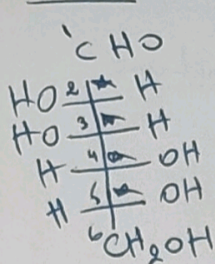
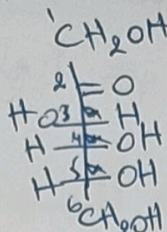


EXO 1: 11,5 pts



A 0,25



B 0,25

$$A = 2^{n-2} = 2^4 = 16$$

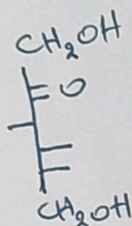
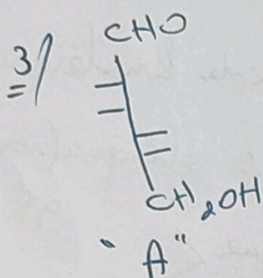
0,5

$$B = 2^{n-3} = 2^3 = 8$$

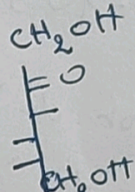
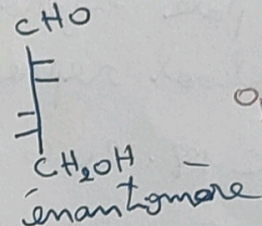
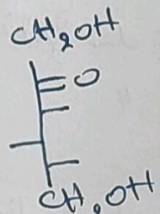
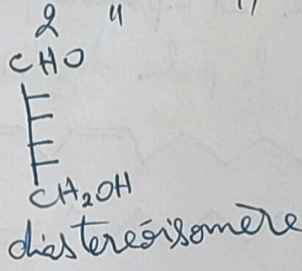
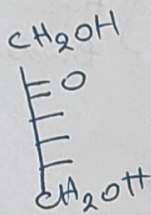
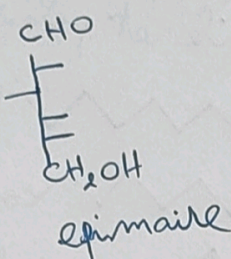
0,5

2/A = 1 fct aldéhyde "C_n"
4 fcts alcool secondaires "C₂, C₃, C₄ et C₅"
1 fct alcool primaire "C₆".

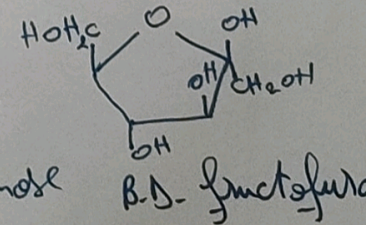
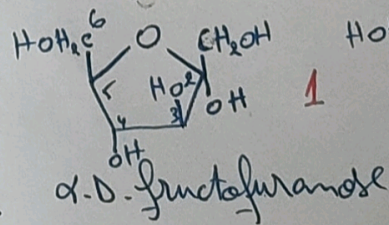
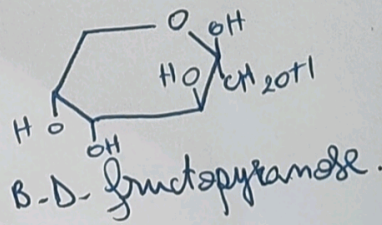
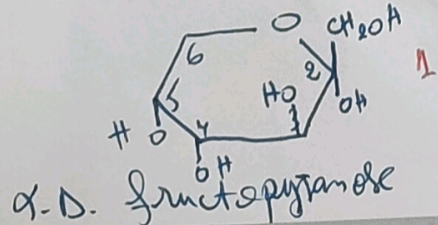
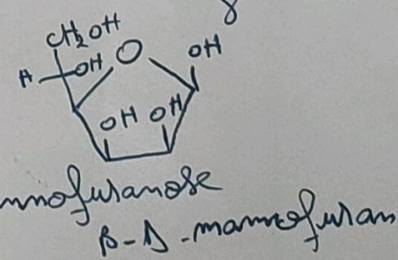
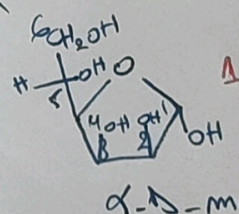
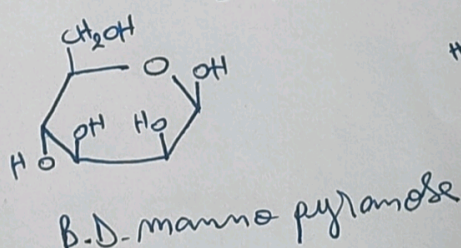
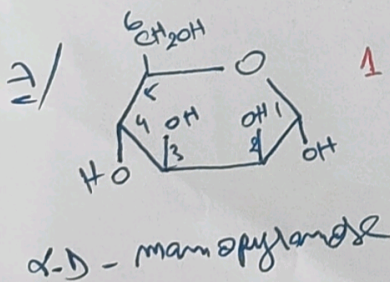
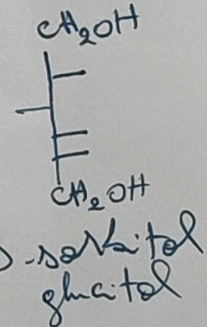
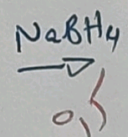
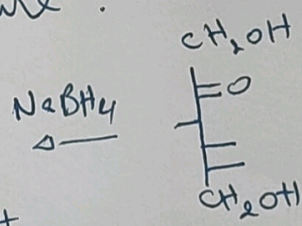
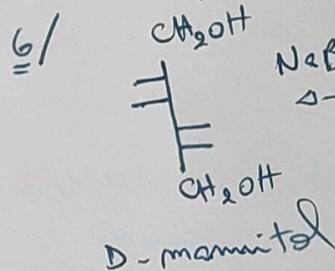
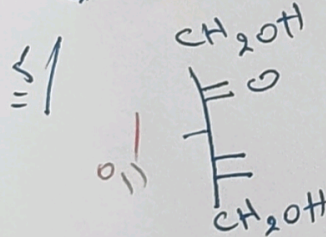
B = 1 fct cétone "C₂"
3 fcts alcool secondaires "C₃, C₄ et C₅"
2 " " primaires "C₁ et C₆"



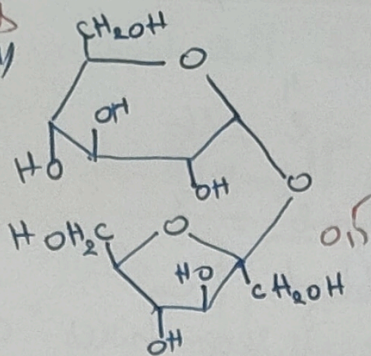
"B"



4/ La présence d'un grand nbr d'atomes d'oxygène très électronégatifs
ce qui les liaisons chimiques sont polarisées ce qui entraîne la
création d'une molécule polaire.



EX02 : 2 pts = 11



2 / mon réducteur, il n'a pas un OH
hémiacétalique libre. 1

3/ liaison osidique "glycosidique"

Exo 3: 6h pta

EX03: 6,1 pts

1/ phosphatidyl-sérine. Il n'est pas soluble dans l'eau parce qu'il a un caractère amphipathique, la partie de l'acide gras représente le pôle hydrophobe et apolaire, le groupement phosphate avec l'alcool représente le pôle hydrophile et polaire.

l'alcool représente le pole hydrophile et l'acide gras le pole hydrophobe

2 1 C18: 2w6 1  Acide linoléique

C18: 1w9 1  Acide oléique (monosaturé)

nbr de "C" $\nearrow$ point de fusion $\nearrow$ Acide linoléique $\swarrow$
 nbr de "=" $\nearrow$ " $\searrow$ Acide stéarique $\searrow$
 0,1 0,1

3/ sérine + acide phosphatidique. 1