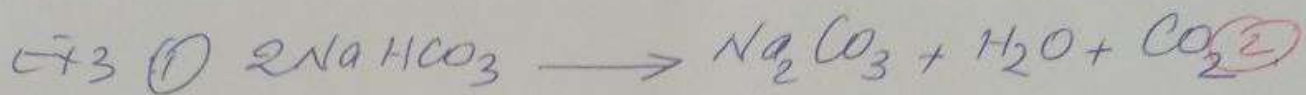


correction type 1

$$EX_1: 2+1+1+1$$

EX1 et EX2 = voir le cours EX2: 1+2



perte totale: 36,6% perte $CO_2 = 25,4\%$

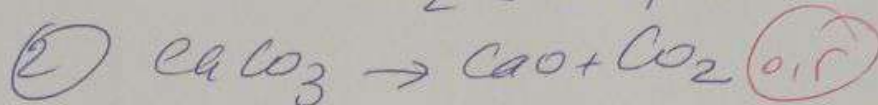
\Rightarrow perte $H_2O = 11,2\%$

on fait le calcul théorique:

perte totale = 36,9% $\textcircled{0,15}$

" $H_2O = 10,71\%$ $\textcircled{0,1}$

" $CO_2 = 26,19\%$ $\textcircled{0,25}$



le CO_2 est issu uniquement de $CaCO_3$.

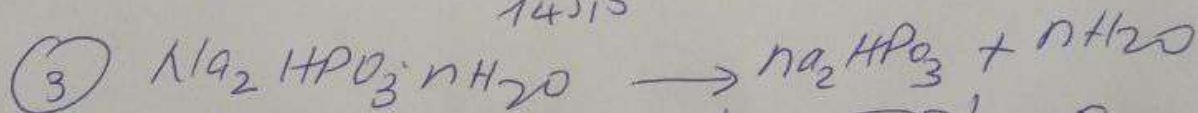
donc $n_{CO_2} = n_{CaCO_3}$

$$n_{CO_2} = (145,3 - 115,4) / 44 = 0,682 \text{ mmole}$$

donc $n_{CaCO_3} = 0,682 \text{ mmole} \Rightarrow m = 0,682 \times 100 \text{ g} = 68,2 \text{ mg}$

$$m_{CaCO_3} = 68,2 \text{ mg}$$

$$\% CaCO_3 = \frac{68,2}{145,3} = 46,9\% \textcircled{1}$$



$$126 + 18n$$

$$250$$

$$126$$

$$145,7$$

$$+ 18n$$

$$\textcircled{2}$$

$$\Rightarrow n = 5$$

EX4: on fait le calcul nécessaire $\textcircled{1}$

$$K = 0,2814, N_{\text{mano}} = 883,7 \text{ } \mu\text{mole}$$

$$S_{\text{BET}} = 86,8 \text{ m}^2/\text{g}$$

$\textcircled{1}$